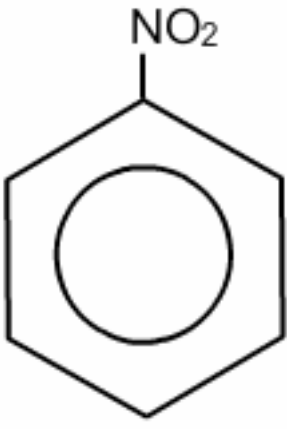
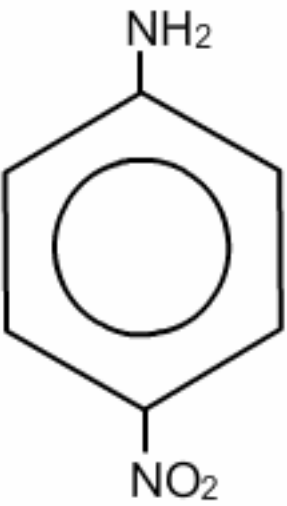
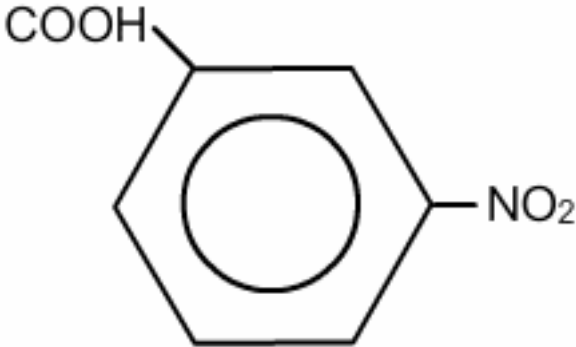
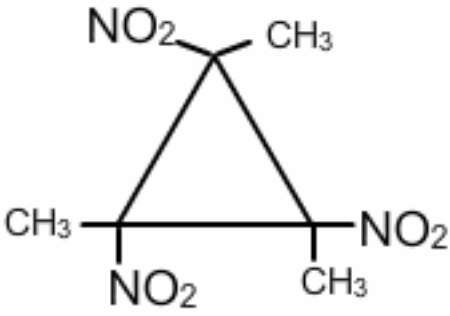
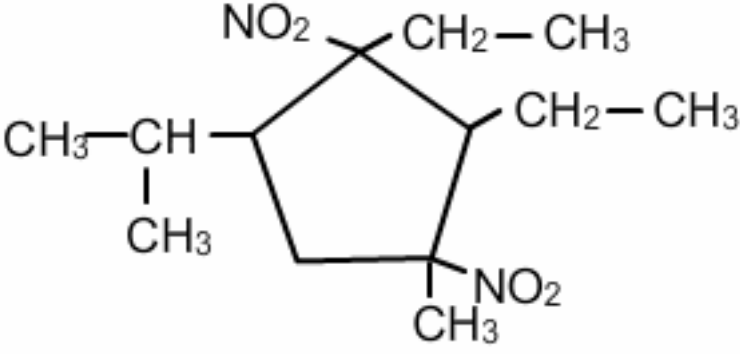
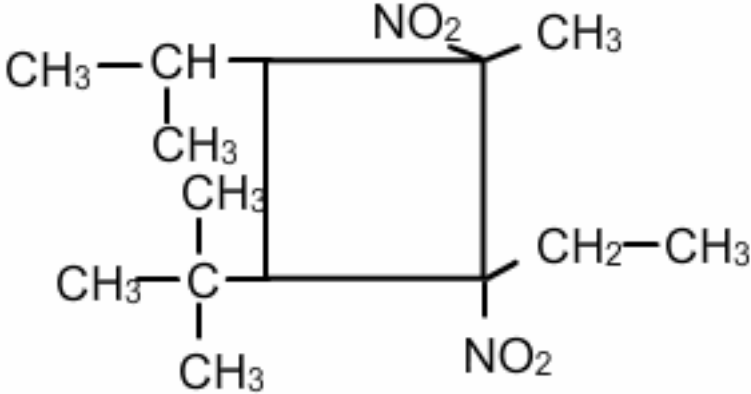
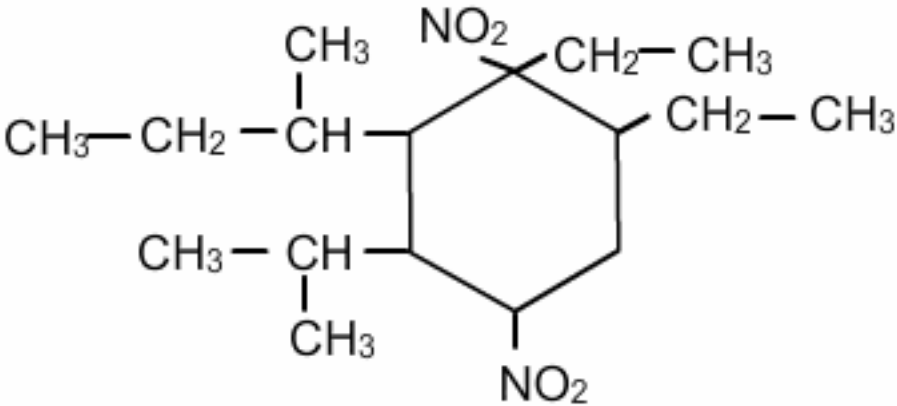


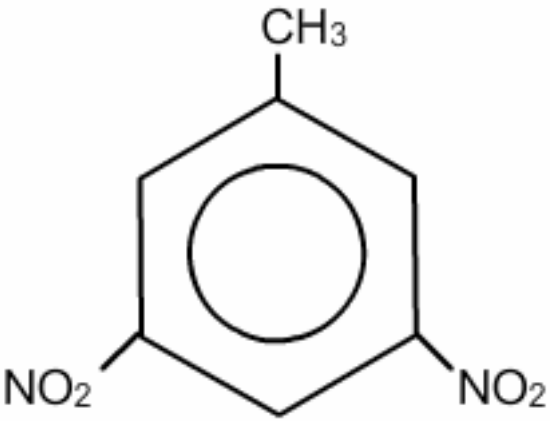
EJERCICIOS NOMENCLATURA DE ÁCIDOS CARBOXÍLICOS

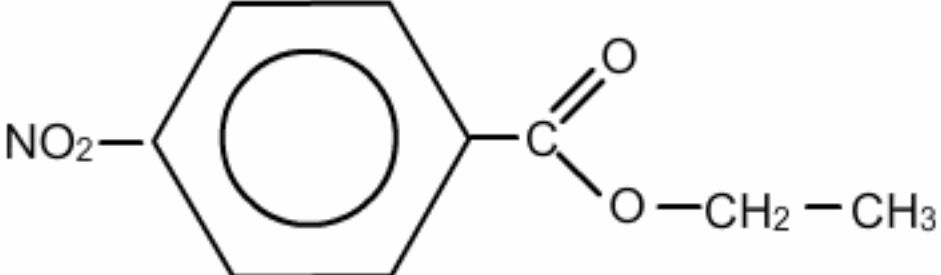
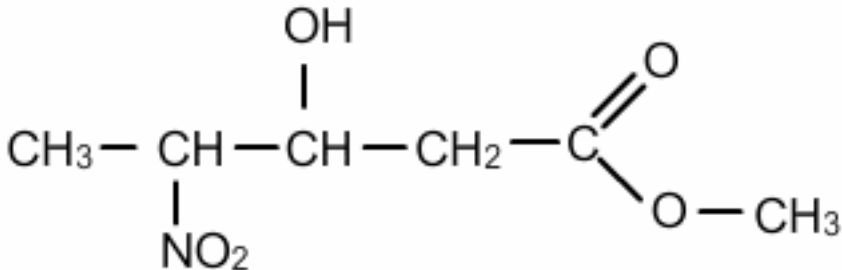
N°	Fórmula	Nombre
1	$\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{NO}_2$	
2		
3	$\begin{array}{ccccccc} \text{CH}_3 & - & \text{CH} & - & \text{CH} & - & \text{CH} & - & \text{CH}_3 \\ & & & & & & & & \\ & & \text{NO}_2 & & \text{CH}_3 & & \text{NO}_2 & & \end{array}$	
4		
5	$\begin{array}{ccccccc} & & & & & & \text{CH}_3 \\ & & & & & & \\ \text{CH}_2 & - & \text{C} = & \text{CH} & - & \text{CH} & \\ & & & & & & \\ \text{NO}_2 & & \text{NO}_2 & & \text{NO}_2 & & \end{array}$	

6	$\begin{array}{ccccccc} & & & \text{CH}_3 & & & \\ & & & & & & \\ \text{CH}_2 & - & \text{C} = & \text{CH} & - & \text{C} & - \text{CH}_2 - \text{NO}_2 \\ & & & & & & \\ \text{NO}_2 & & \text{CH}_3 & & & \text{NO}_2 & \end{array}$	
7		
8	$\begin{array}{ccccccc} & & & \text{CH}_3 & & & \\ & & & & & & \\ \text{CH}_2 & - & \text{CH} & - & \text{CH}_2 & - & \text{C} & - \text{CH}_2 - \text{NO}_2 \\ & & & & & & \\ \text{NO}_2 & & \text{CH}_3 & & & \text{NO}_2 & \end{array}$	
9	$\begin{array}{ccccccc} & & \text{O} & & \text{NO}_2 & & \text{O} \\ & & & & & & \\ \text{CH}_3 & - & \text{C} & - & \text{C} = & \text{C} & - & \text{C} \\ & & & & & & & \backslash \\ & & & & & \text{NO}_2 & & \text{H} \end{array}$	
10	$\begin{array}{ccccccc} & & \text{O} & & & & \text{O} \\ & & & & & & \\ \text{CH} \equiv \text{C} & - & \text{C} & - & \text{C} & - & \text{CH} & - & \text{C} \equiv \text{C} & - & \text{C} \\ & & & & & & & & & & \backslash \\ & & \text{CH}_2 & & \text{NO}_2 & & & & & & \text{H} \end{array}$	
11	$\begin{array}{ccccccc} & & \text{O} & & & & \text{O} \\ & & & & & & \\ \text{CH}_3 & - & \text{C} & - & \text{CH} & - & \text{CH}_2 & - & \text{C} \\ & & & & & & & & \backslash \\ & & & & \text{NO}_2 & & & & \text{OH} \end{array}$	

12	$\begin{array}{ccccccc} & & \text{CH}_2 - \text{CH}_3 & & & \text{O} & \\ & & & & & & \\ \text{NO}_2 - & \text{CH} & - \text{CH} & - \text{CH} & - \text{C} & - \text{NO}_2 \\ & & & & & & \\ & & \text{NO}_2 & & \text{OH} & & \end{array}$	
13	$\begin{array}{ccccccc} & & \text{O} & & & & \\ & & & & & & \\ \text{CH}_3 - & \text{C} & - \text{CH} & - \text{CH}_3 \\ & & & & & & \\ & & \text{NO}_2 & & & & \end{array}$	
14	$\begin{array}{ccccccc} & & & & \text{O} & & \\ & & & & & & \\ \text{CH} \equiv \text{C} - & \text{C} & = \text{C} & - \text{C} & - \text{CH}_3 \\ & & & & & & \\ & \text{OH} & \text{NO}_2 & & & & \end{array}$	
15	$\begin{array}{ccccccc} & \text{NO}_2 & \text{NO}_2 & & \text{O} & & \\ & & & & & & \\ \text{CH}_3 - & \text{C} & - \text{CH} & - \text{C} & - \text{C} & \begin{array}{l} \text{O} \\ // \\ \text{H} \end{array} \\ & & & & & & \\ & \text{O} - \text{CH}_2 - \text{CH}_3 & & & & & \end{array}$	
16	 <p>A cyclopropane ring with four substituents: a methyl group (CH₃) and a nitro group (NO₂) on each of the two carbons at the top and bottom of the ring.</p>	
17	 <p>A cyclopentane ring with five substituents: a methyl group (CH₃) and a nitro group (NO₂) on the left carbon; an ethyl group (CH₂-CH₃) and a nitro group (NO₂) on the top carbon; and a methyl group (CH₃) and an ethyl group (CH₂-CH₃) on the right carbon.</p>	

18	$\begin{array}{c} \text{CH}_2 - \text{CH} = \text{CH} - \text{CH} - \text{C} \\ \qquad \qquad \qquad \qquad \qquad \qquad // \\ \text{NO}_2 \qquad \qquad \qquad \text{CH}_3 \qquad \qquad \text{O} \\ \qquad \qquad \qquad \qquad \qquad \qquad \qquad \qquad \backslash \\ \qquad \qquad \qquad \qquad \qquad \qquad \qquad \qquad \text{OH} \end{array}$	
19	 <p>The structure shows a cyclobutane ring with the following substituents: a methyl group (CH₃) on the top-left carbon, a nitro group (NO₂) on the top-right carbon, a methyl group (CH₃) on the bottom-right carbon, and an ethyl group (CH₂-CH₃) on the bottom-left carbon. The bottom-left carbon also has a methyl group (CH₃) attached to it.</p>	
20	$\text{CH}_2 = \text{CH} - \underset{\text{NO}_2}{\text{CH}} - \underset{\text{OH}}{\overset{\text{O}}{\text{C}}}$	
21	$\text{CH}_3 - \underset{\text{OH}}{\text{C}} = \text{CH} - \overset{\text{O}}{\text{C}} - \text{CH} = \text{CH}_2 - \text{NO}_2$	
22	 <p>The structure shows a cyclohexane ring with the following substituents: a methyl group (CH₃) on the top-left carbon, a nitro group (NO₂) on the top-right carbon, an ethyl group (CH₂-CH₃) on the right carbon, a methyl group (CH₃) on the bottom-right carbon, a nitro group (NO₂) on the bottom carbon, and an ethyl group (CH₂-CH₃) on the left carbon.</p>	

23	$\text{CH} \equiv \text{C} - \underset{\substack{ \\ \text{OH}}}{\text{C}} = \underset{\substack{ \\ \text{NO}_2}}{\text{C}} - \overset{\text{O}}{\parallel}{\text{C}} - \underset{\substack{\text{O} \\ \parallel \\ \text{OH}}}{\text{C}}$	
24	$\text{NO}_2 - \underset{\substack{ \\ \text{OH}}}{\text{C}} = \overset{\substack{\text{CH}_3 \\ }}{\text{C}} - \underset{\substack{ \\ \text{CH}_2 - \text{CH}_3}}{\text{C}} = \overset{\substack{\text{CH}_3 \\ }}{\text{C}} - \overset{\substack{\text{CH}_2 \\ }}{\text{C}} - \overset{\text{O}}{\parallel}{\text{C}} - \text{NO}_2$	
25	$\text{NO}_2 - \overset{\substack{\text{CH}_3 \\ \\ \text{CH}_2 \\ }}{\text{CH}} - \underset{\substack{ \\ \text{OH}}}{\text{C}} = \text{CH} - \overset{\text{O}}{\parallel}{\text{C}} - \underset{\substack{ \\ \text{NO}_2}}{\text{CH}} - \text{C} \equiv \text{CH}$	
26		
27	$\text{NO}_2 - \underset{\substack{ \\ \text{CH}_3}}{\text{CH}} - \overset{\substack{\text{OH} \\ }}{\text{CH}} - \text{CH}_2 - \overset{\text{O}}{\parallel}{\text{C}} - \text{O} - \text{CH}_2 - \text{CH}_3$	

28		
29		
30	